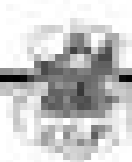


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Exam Practice Answer Book

Lesson 2: Exploring Group 1

Lesson overview

OCR Specification reference

OCR G4 1

Learning objectives

- Explore the properties of Group 1 metals.
- Compare their reactivity.
- Relate their reactivity to their electronic structures.

Learning outcomes

- Explain why Group 1 metals are known as the alkali metals (21).
- Use the trends down the group to make predictions (22).
- Relate the properties of alkali metals to their electron configurations (23).

Skills development

- WS 2.4 Identify the main hazards associated with Group 1 metals.
- WS 2.8 Make and record observations.
- WS 4.1 Use scientific vocabulary, terminology and definitions.

Maths facts: Interpolate and extrapolate graphs

Resources needed: Equipment as listed in the Technician's notes, Worksheets 4.2.1, 4.2.2 and 4.2.3, Technician's notes 4.2

Digital resources: Presentation 4.2 (10)

Key vocabulary: alkali, density, diffusion, ion, reactivity, stable electronic structure

Teaching and learning

Engage

- Show the bottles of Group 1 metals (Technician's notes 4.2) and ask students: what about why they are stored under oil and why the potassium being used are so small (21)?

Challenge and develop

- Demonstrate the reactions of lithium, sodium and potassium with water following the method detailed in Technician's notes 4.2. Encourage the prediction between them: the way they become softer and less dense going down the group and the way that reactivity increases going down the group (22).
- Follow up the demonstration with a video clip so students can see close-up of the metals and their reactions with water (through search string) (Must include in notes) (23)

Explain

Low demand

- Students read Properties of Group 1 elements and Reaction trends of alkali metals on pages 102-103 in the Student Book and answer questions 1-5 (21)

Standard and high demand

- Students read the prediction and differences between the reactions of lithium, sodium and potassium with water (21)

Consolidate and apply

Low and standard demand

- Students use the trends down the group to make predictions using Worksheets 4.2.1 and 4.2.2 (22)
- Students complete Worksheet 4.2.3 (23)

High demand

- Students use the trends in melting point down the group to make predictions using Worksheet 4.2.2 (22)
- Point out fluorine at the bottom of Group 7. Explain that it is very rare and that all its isotopes are radioactive, so it is nearly impossible to study. Ask students to predict how it reacts with water and to justify their predictions using evidence from reactions they have seen (24)
- Show Presentation 4.2.3 and discuss the similarity between the electronic structures of lithium and sodium (24)
- Put pairs or small groups to explain how the electronic structures of these elements make them have similar chemical properties (24)

Extend

Ask students able to progress further to research the reactions of hydrogen, magnesium and calcium with water, compare them with Group 1 elements and describe any similarities and differences they notice

Plenary inspiration

Ask students to draw a large triangle with a smaller inverted triangle that just fits inside it (see they have that example). In the next three lessons to write something they've seen, something they've done, and something they've discussed. Then add something they've learned in the next 3 lessons.

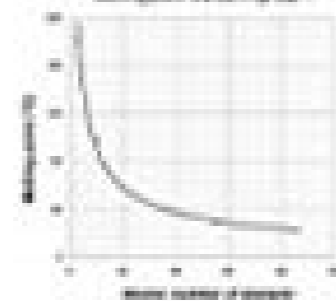
Answers to Worksheet 4.2.1

- They will react more readily with water than potassium because the reactivity of Group 1 metals increases going down the group: $2\text{Li} + 2\text{H}_2\text{O} \rightarrow 2\text{LiOH} + \text{H}_2$, $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$, $2\text{K} + 2\text{H}_2\text{O} \rightarrow 2\text{KOH} + \text{H}_2$, Calcium hydroxide, Hydrogen, 4. The hydroxide dissolves in water to make alkaline solutions

Answers to Worksheet 4.2.2

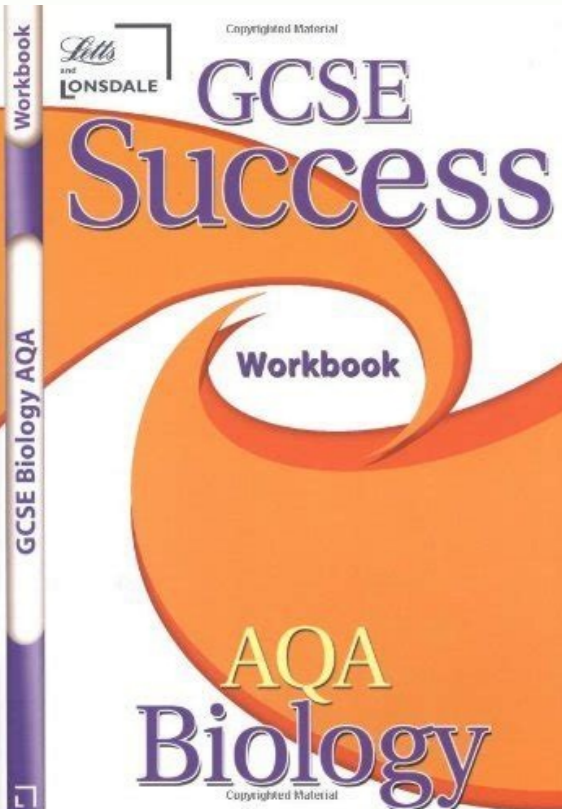
Using graphs to make predictions: potassium, melting point 63-64 °C, francium, melting point 24-26 °C

Reactivity trends in group 1



Answers to Worksheet 4.2.3

- They both form the potassium ion in their outer shell: K^+ . In one of the graphs changes in the number is to large compared to K^+ .



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